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3,4,5. ionic compound. formula unit. lattice energy. polyatomic ion. A compound that consists of positive and negative ions. the lowest whole-number ratio of ions in an ionic compound.

Chapter 6 Section Quiz Introduction To Chemical Bonding

Chemical Bonding CHAPTER 6 Section 1

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Introduction to Chemical Bonding What is a chemical bond and why does it form? Section 2 Covalent Bonding and Molecular Compounds What is a molecular formula? What are the characteristics of a covalent bond? How does a Lewis structure represent a molecule?

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Chapter 6 CHEMICAL NAMES AND FORMULAS

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Bonding SECTION 1 SHORT ANSWER

Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis

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2 Covalent Answer Key ...

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REPRESENTING CHEMICAL COMPOUNDS
Section 6.3 IONIC CHARGES Section 6.4
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CHAPTER 6 REVIEW Chemical Bonding
SECTION 3 SHORT ANSWER Answer the

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following questions in the space provided.

1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom.

2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions.

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Chapter 6 Section 3 Chemical Bonding

Chapter 6: Chemical Bonding Section 1- Introduction to Chemical Bonding

Objectives: define chemical bond; differentiate between covalent and ionic bonding; explain why bonding occurs; use the difference in electronegativity to determine whether a bond is polar

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covalent, nonpolar covalent or ionic

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Section 6.2 – Covalent Bonding. A covalent bond is a chemical bond in which two atoms share a pair of valence electrons. When two atoms share one pair of electrons, the bond is called a

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Modern Chemistry 9 Chemical Bonding
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Bonding SECTION 3 IONIC BONDING AND
IONIC COMPOUNDS SHORT ANSWER
Answer the following questions in the
space provided. 1. _____ The notation for

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sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal.

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youmakemestrongmm. BUSI 211: Week
1 - Focus Points 11 Terms.

Chapter 6 - Chemical Bonds

Chemical Bonding: Chapter 6 - Section 2.
Covalent Bonding and Molecular

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Compounds. STUDY. PLAY. Diatomic molecule. a molecule containing only two atoms. Anion. a negative ion. Cation. a positive ion. Molecules. a neutral group of atoms that are held together by covalent bonds. Molecular compound.

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Answer the following questions in the space provided

1 a The notation for sodium chloride, NaCl, stands for one (a) formula unit (c) crystal (b) molecule (d) atom

2 d In a crystal of an ionic compound, each cation is ...

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CHAPTER 6 REVIEW Chemical Bonding

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following questions in the space

provided. 1. a A chemical bond between atoms results from the attraction

between the valence electrons and of different atoms. (a) nuclei (c) isotopes

(b) inner electrons (d) Lewis structures

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2. b A covalent bond consists of (a) a shared electron.

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CHAPTER 6 REVIEW Chemical Bonding

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SECTION 6-2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms. 2. Name two elements that form compounds that are exceptions to the octet rule.

Chapter 6 Section 6-1 Review

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