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SECTION 9.2. PROBLEMS Write the
answer on the line to the left. Show all
your work in the space provided. 1. The
following equation represents a
laboratory preparation for oxygen gas:

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Write the answer on the line to the left.
Show all your work in the space
provided. 1. The following equation
represents a laboratory preparation for
oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$... CHAPTER 9
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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to

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Reactions SECTION 82 SHORT ANSWER
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From above we can see that if we have

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12.4 mol H₂ we need 4.13 mol N₂. We don't have that much N₂ so the .892 mol of N₂ must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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Section

CHAPTER 9 REVIEW Stoichiometry
SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. b The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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CHEMISTRY NOTES - Chapter 9 Stoichiometry

Chapter 9 - Stoichiometry Review #1 - #18, #31, & #38 Answers . 38. To ensure that all magnesium is converted to MgO, I would use pure oxygen, not air, to carry out the reaction, because Mg could react with N₂ in air to form

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MgJN₂. The pure oxygen should be in excess. 5. a. 50 mol HI 6. a. 15.8

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Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of

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reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

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the following questions in the space
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chemical equation represent the (a
masses in grams of all reactants and
products.

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provided. 1. _____ The coefficients in a
chemical equation represent the (a)
masses in grams of all reactants and
products. (b) relative number of moles of

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reactants and products.

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SECTION 9-3 PROBLEMS Write the
answer on the line to the left. Show all
your work in the space provided. 1. 88%
If the actual yield of a reaction is 22 g

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and the theoretical yield is 25 g,
calculate the percent yield. 2. 6.0 mol of
N₂ are mixed with 12.0 mol of H₂
according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g)
N₂; 2.0 mol a.

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Chapter 9 Stoichiometry Definition •

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Stoichiometry -Relationship between quantities • Composition stoichiometry -The mass relationships of elements in compounds (Ch 7.3) • Reaction stoichiometry -The mass relationships between reactants and products in a chemical reaction

Section 1 Introduction to Stoichiometry

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