

## Chem 1411 Chapter 6 Thermochemistry Exercises

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6 Thermochemistry 228 6.1 The Nature of Energy 229 6.2 Enthalpy and Calorimetry 235 CHEMICAL IMPACT Nature Has Hot Plants 238 ... (in Chapter 6) because of the importance of energy in various chemical processes and models, but the ... J. Chem. Ed. 68 (1991): 400. 1.5 Significant Figures and Calculations 15 Note that for multiplication and ...

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Chapter 12 Question 7C In a study of the decomposition of nitrous oxide at 565degree C N<sub>2</sub>Og rig Chapter 13 Question 6 Consider the reaction Cs 122gCOg Write the equilibrium constant f Chapter 13 Question 8C The equilibrium constant K<sub>p</sub> for the following reaction is 2.74 at 1150 K A chemist starts with 50.0 mL of a 0.415 M NaCl solution.

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1 The properties of gases 1A The perfect gas Answers to discussion questions 1A.2 The partial pressure of a gas in a mixture of gases is the pressure the gas would exert if it occupied alone the same container as the mixture at the same temperature.

### **Chem 1411 Chapter 6 Thermochemistry**

March 14 chapter 5 practice multiple choice questions and answers Chapter 5 Worked-Examples Clickers questions Ch 5 Chapter 6: Electronic Structure of Atoms Clickers questions Ch 6 HOMEWORK 6\_Due. March 28 chapter 6 practice multiple choice questions with answers worked Examples Ch 6 Practice Exam - CHEM 1311 EX 2 fall-2018- chapters 4-6 ...

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Chapter 1. Practice Exercises 1.1 (a) SF<sub>6</sub> contains 1 S and 6 F atoms per molecule (b) (C<sub>2</sub>H<sub>5</sub>)<sub>2</sub>N<sub>2</sub>H<sub>2</sub> contains 4 C, 12 H, and 2 N per molecule (c) Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> contains 3 Ca, 2 P, and 8 O atoms per formula unit (d) Co(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O contains 1 Co, 2 N, 12 O, and 12 H per formula unit. 1.2 (a) NH<sub>4</sub>NO<sub>3</sub> contains 2 N nitrogen, 4 H hydrogen, 3 O oxygen atoms per formula unit (b) FeNH<sub>4</sub>(SO<sub>4</sub>)<sub>2</sub> contains 1 Fe iron ...

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D) 49.6 kJ MST You have 10.5 moles of Al. If you multiply by the atomic mass of Al, you get the grams of Al. The amount of energy needed is the grams of Al heat capacity temperature change. 10.5 moles x 27al=283.5g (283.5 x 0.900)(30.5-225)= 4926 4926 x 1/1000= 49.6

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