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Mole 1.. Idenü\$-' and calculatethe number of.representa- five particles. in each of the following quantities. ...

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Chapter 10 - The Mole Section 10.1  
Measuring Matter 1. pair 2. 5 3. dozen 4.  
gross 5. 200 6. ream 7. 6,000,000,000 8.  
0.5 mol 9. 6.02  $\times 10^{23}$  10. four moles 11.  
6.02  $\times 10^{23}$  Cu atoms

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Likewise, a mole will always contain  $6.022 \times 10^{23}$  particles. Although a mole of any substance is the same number, different substances have different molar masses. The molar mass of a compound can be calculated from its chemical formula and can be used to convert from mass to moles of that compound.

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a Mole? Avogadro's Number Molar Mass  
Mole Conversions Mass to Moles  
Particles to Moles Volume to Moles  
Empirical and Molecular Formulas %  
Composition Determining Empirical  
Formula Determining Molecular Formula

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320 Chapter 10 • The Mole Section  
110.10.1 Measuring Matter MAIN Idea  
Chemists use the mole to count atoms,  
molecules, ions, and formula units. Real-  
World Reading Link Has your class ever  
had a contest to guess how

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10 161 Section 10.1 Measuring Matter  
page 320–324 Practice Problems pages  
323–324 1. Zinc (Zn) is used to form a  
corrosion-inhibiting surface on  
galvanized steel. Determine the number  
of Zn atoms in 2.50 mol of Zn. 2.50 mol  
Zn

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Chapter 10 The Mole 1. The Mole 2. Not  
the type of mole we are talking about ...  
Chemistry - Chp 12 - Stoichiometry -  
PowerPoint Mr. Walajtys. The mole  
concept Kris Ann Mae Yap Bonilla.  
Chapter 3 Chemical Formulae and  
Equations M BR. Chapter 10: The Mole  
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