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cm has only two significant figures.
When adding and subtracting, limit and
round your answer to the least number
of decimal ... Chemistry Worksheet:
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USING SIGNIFICANT FIGURES ... any of the factors. Example 1: $23.0 \text{ cm} \times 432 \text{ cm} \times 19 \text{ cm} = 188,784 \text{ cm}^3$ The answer is expressed as $190,000 \text{ cm}^3$ since 19 cm has only two significant figures. When adding and subtracting, limit and round your answer to the least number of decimal ... Chemistry Worksheet: Significant ...

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The final answer in a multiplication or division problem should contain the same number of significant figures as the original number with the fewest significant figures. In addition and subtraction, the final answer should contain the same number of decimal places as the original number with the fewest number of decimal places.

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Sig Fig pg 1

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answers page 70 - the mole and volume
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Worksheet 1 - Scientific notation/ significant figures

Significant Figures Worksheet I)

Determine the number of significant
figures in each of the following:

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75.02mm 18.90 ml. 12 test tubes 0.0049
g 150 cm 150. Cm 2) Re-write the
quantity 827,000,000,000,000
picoseconds to show: a) 1 sig. fig. b) 2
sig. figs. c) 3 sig. figs. d) 4 sig. figs. e) 5
sig. figs. B

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work answers chemistry if8766,
Chemistry if8766 46 naming molecular
compounds, Answers combined gas law,
, Trom po no, , Significant figures name.

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Significant figures worksheet answers

measurements in scientific notation,
which has two parts. N of 10 by which
only one digit is 10^n A number in which
placed to the left of the decimal An
exponent the number is multiplied Thus,
the temperature of the Sun, 15 million

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kelvins, is written as 1.5×10^7 K in scientific notation. Positive Exponents Express 1234.56 in scientific notation.

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Significant Figures Practice

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Scientific Notation Use with Appendix B, Scientific Notation

Calculations using Significant Figures
When multiplying and dividing, limit and round to the least number of significant figures for any of the factors. Example 1:
 $23.0 \text{ cm} \times 432 \text{ cm} \times 19 \text{ cm} = 188,784$

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cm 3 The answer is expressed as 190,000 cm 3 since 19 cm has only two significant figures.

Calculations using Significant Figures - Mr. V's Chemistry ...

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includes: page 9 - significant figures
page 12 - temperature and its
measurement page 13 - freezing and
boiling point graph

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Chemistry IF8766 9 OInstructional Fair,
Inc. 59. CALCULATIONS USING
SIGNIFICANT FIGURES ... number of
significant figures in any of the factors.
11 Example 1 : $23.0 \text{ cm} \times 432 \text{ cm} \times 19$
 $\text{cm} = 188,784 \text{ cm}^3$ 11 The answer is
expressed as $190,000 \text{ cm}^3$ since 19 cm
has only two significant fiaures.

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Zeros at the end of a number and to the right of a decimal are significant: 85.00 g has four sig figs; 9.000 000 000 mm has 10 sig figs. Zeros at the end of a number but to the left of a decimal may or may not be significant. If such a zero has been measured, or is the first estimated digit, it is significant.

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Department of Chemistry and Physics:
Worksheet 1: Scientific

Notation/Significant Figures. 1. Convert each of the following into scientific notation. ... Give the answer in correct scientific notation. a) $3.95 \times 10^2 / 1.5 \times 10^6$ b) $(3.5 \times 10^2)(6.45 \dots$ Scientific

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Notation/ Significant Figures Worksheet
2