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### **24 ChemQuest 8 Name: - Yumpu**

24 ChemQuest 29 Monc LcW s Stnuctunes tr l Name Date: Hour: Introduction Questions Draw the Lewis structures for the following molecules. Part c is done for you. a) H<sub>2</sub>O (oxygen is central) b) SO<sub>2</sub> c) N<sub>2</sub> It t-t (a [J]rt @ N=N @ @ f\ @ 2 In question 1 above you should see three different kinds of bonds of bonds as best you can. a) single bond: one es ...

### **Chemquest 24 More Lewis Structures**

24 ChemQuest 8 Name: \_\_\_\_\_ Date: \_\_\_\_\_ Hour: \_\_\_\_\_ Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

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Note: the only difference in the resonance structures is the position of the double bonds. Critical Thinking Questions 9. a) The nitrate ion, NO<sub>3</sub><sup>-</sup>, is best represented by three resonance structures. Draw the three structures below. Note: the only difference between the three structures you draw will be the position of the double bond.

### **Worksheet: More Lewis Structures Name**

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### **Bonding | mretter**

Atomic Structure - Part 2. ChemQuest #24: Ionic Bonding ChemQuest #25: Advanced Ionic Bonding ChemQuest #26: Transition Metals Read: ... ChemQuest #28: Lewis Structures ChemQuest #29: More Lewis Structures ChemQuest #31: Molecular Geometry ChemQuest #32: Electronegativity Read:

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### **Information : Drawing Covalent Compounds**

Worksheet: More Lewis Structures Name \_\_\_\_\_ CHEMISTRY: A Study of Matter © 2004, GPB 5.14 I. Draw Lewis Structures for the following molecules.

### **ChemQuest 23 : Drawing Covalent Compounds**

Information: Steps for Drawing Lewis Structures for Covalent Compounds Study the two examples in the table of how to write structures for CO<sub>3</sub><sup>2-</sup> and NH<sub>3</sub>. Make sure you understand each of the five steps. CO<sub>3</sub><sup>2-</sup>-NH<sub>3</sub> Step #1: Add up the number of valence electrons that should be included in the Lewis Structure. 4 + 3(6) + 2 = 24 (carbon has four ...

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Therefore, the charge on each X is 1. Here's why: Remember Answer: + that S has a negative 2 charge. There are two "X" ions to balance out the negative 2 charge of S.

**Atomic Structure - Part 2 - CHM 1032**

Information: Steps for Drawing Lewis Structures for Covalent Compounds Study the two examples in the table of how to write structures for CO<sub>3</sub><sup>2-</sup> and NH<sub>3</sub>. Make sure you understand each of the five steps. CO<sub>3</sub><sup>2-</sup>-NH<sub>3</sub>  
Step #1: Add up the number of valence electrons that should be included in the Lewis Structure. 4 + 3(6) + 2 = 24 (carbon has four ...

**63 ChemQuest 18 Electron Configurations66 ChemQuest 19 ...**

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