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Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N

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Stoichiometry Molar Mass The trick: • By definition, this is the mass of 1 mol of a substance (i.e., g/mol) – The molar mass of an element is the mass number for the element that we find on the periodic table – The formula weight (in amu's) will be the same number as the molar mass (in g/mol)

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The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

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## **VIBRATIONS AND WAVES**

Stoichiometry Study Guide KEY Chemistry RHS Mr. Moss The Ultimate Guide to Stoichiometry for AP Chemistry Stoichiometry is the method of quantitatively relating the changes in substances undergoing a chemical reaction. This idea is present in almost every topic in AP Chemistry, so it is one of the most important areas to study for your exam.

## **mc06se cFMsr i-vi**

View Notes - note\_taking\_guide\_801 from CHEM 303 at Wheaton College. Note Taking Guide: Episode 801 Stoichiometry study of the quantity based on balanced Name relationships in a chemical

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Key Words to Know: (Check off the ones that you know and study the other ones!) Vocab Definitions/Related Words Topic 1: Reviewing Compounds o Conservation of Mass o Reactant vs. Product

### **Chapter 3 Stoichiometry - Chemistry**

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