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Explain your answer by referring to the diagram. 5. A list of compounds and elements is given

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CO-APO 6. How can
you tell the difference
between a

homogeneous mixture
and a heterogeneous

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Chapter 3 ...

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Chapter 3 61 1. Into

what two broad classes

can all matter be

divided? 2. What is the

difference between a

mixture and a pure

substance? 3. What is

the difference between

a compound and an

element? 4. Can a

compound be a

heterogeneous

mixture? Explain your

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Ions 1. What are the
names of the two
elements shown? Date
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2. Are the elements

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Chapter 7, Section 7.2

1. How many valence
electrons does a
neutral magnesium
(Mg) atom have? 2.

What is the charge on

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a magnesium ion?

What does magnesium
have to do to form
such ... Teaching

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or Insoluble? Use your

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solubility table to determine if the following chemicals are soluble or insoluble in water. Write the terms "aqueous" or "solid" next to each chemical.

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Chapter 6 35 1. What

are the four sections,

or blocks, of the

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periodic table? 2. What does each block represent? 3. What do elements in the s-block have in common? 4. What is the valence electron configuration of each element in group 1? 5.

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the shape of an s

orbital? 2. What is the

relationship between

the size of an s orbital

and the principal

energy level in which it

is found? 3. What is the

shape of a p orbital?

How many p orbitals

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are there in a sublevel?

4. How many electrons
can each ...

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shape of an s orbital?

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